
1. What is the oxidation number of manganese in KMnO_4 ?

- A) +7 B) +2 C) +3 D) +4

2. Which polyatomic ion has a charge of 3-?

- A) chromate ion B) oxalate ion
C) phosphate ion D) thiocyanate ion

3. What is the oxidation state of nitrogen in the compound NH_4Br ?

- A) -1 B) +2 C) -3 D) +4

4. What is the oxidation number of sulfur in $\text{Na}_2\text{S}_2\text{O}_3$?

- A) -1 B) +2 C) +6 D) +4

5. Given the balanced equation representing a reaction:



The oxidation state of chlorine in this reaction changes from

- A) -1 to +1 B) -1 to +5
C) +1 to -1 D) +5 to -1

6. What is the oxidation number of chromium in the chromate ion, CrO_4^{2-} ?

- A) +6 B) +2 C) +3 D) +8

7. What is the oxidation number assigned to manganese in KMnO_4 ?

- A) +7 B) +2 C) +3 D) +4

8. What is the oxidation state of nitrogen in NaNO_2 ?

- A) +1 B) +2 C) +3 D) +4

9. In which compound does chlorine have the highest oxidation number?

- A) NaClO B) NaClO_2
C) NaClO_3 D) NaClO_4

10. In which substance does chlorine have an oxidation number of +1?

- A) Cl_2 B) HCl
C) HClO D) HClO_2

11. What is the oxidation number of chromium in $\text{K}_2\text{Cr}_2\text{O}_7$?

- A) +12 B) +2 C) +3 D) +6

12. What is the oxidation number of hydrogen in CaH_2 ?

- A) +1 B) +2 C) -1 D) -2
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