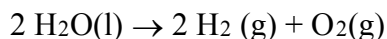
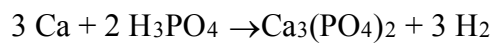


1. Which statement is correct when 18 g of oxygen O₂ reacts with 2.0 g of hydrogen H₂ to form water?



- A) Some oxygen is left over.
B) Water, H₂O, has a molar mass of 20.
C) The Law of Multiple Proportions applies.
D) When 18 g of O₂ reacts with 2.0 g of H₂, 20. g of H₂O is produced.
E) There is insufficient hydrogen for any water to form.
2. What mass of H₃PO₄ is needed to completely react with 30. g of Ca?



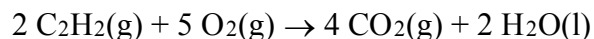
- A) 20 g B) 49 g C) 74 g D) 98 g E) 116 g
3. When 6.000 mole of KClO₃ are reacted, how many grams of KClO₄ are produced?



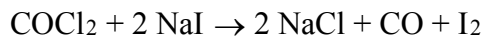
MM	(122.6)	(138.6)	(74.6)
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- A) 104.0 g B) 138.6 g
C) 415.8 g D) 623.7 g
E) 2495 g

4. When 13.0 grams of acetylene, C₂H₂, is reacted, what mass of water, H₂O, is produced?



- A) 9.00 g B) 12.0 g
C) 13.0 g D) 18.0 g
E) 26.0 g
5. In the reaction represented by the equation,



what is the maximum weight of iodine that can be liberated from 60.0 grams of sodium iodide?

- A) 25.4 g B) 50.8 g
C) 102 g D) 153 g
E) 203 g
6. A mixture of 2.0 grams of hydrogen and 32 grams of oxygen is exploded and produces water. What weight of gas remains uncombined ?
- A) 1.0 gram of hydrogen
B) 1.0 gram of oxygen
C) 4.0 grams of hydrogen
D) 8.0 grams of oxygen
E) 16.0 grams of oxygen