**Lab 7: DOUBLE REPLACEMENT REACTIONS AND SOLUBILITY**.

PRELAB PREPARATION

A) For each mixture in the chart:

* 1. Predict the formation of a precipitate in each mixture.
	2. For each mixture write the double replacement reactions, which substances are electrolytes and which are not.
	3. Name each substance in each reaction.
	4. Write the dissociation for each reactant in the chart.

B) Procedure

 1. Obtain eight pipettes and label each with the formulas given on the table.

 2. Clean a micro chemistry plate and shake dry.

 3. **Following the table, mix 4 drops of each**.

 I) Record the presence of a precipitate in the boxes. Record the absence of a precipitate as no reaction.

 ii) Assume a precipitate is any solid produced and/or any sign of turbidity.

 iii) If there is no precipitate, assume no reaction.

|  |  |  |  |  |
| --- | --- | --- | --- | --- |
|  | **AgNO3** | **Ba(NO3)2** | **Ca(NO3)2** | **Mg(NO3)2** |
| **NaOH** |  |  |  |  |
| **Na3PO4** |  |  |  |  |
| **Na2CO3** |  |  |  |  |
| **H2SO4** |  |  |  |  |

**THE NITRATE ION IS ALWAYS SOLUBLE (TABLE F) THEREFORE WE WILL NOT CONSIDER IT A REACTANT IN THE MIXTURES.**

**Na AND H ARE IN GROUP ONE AND ALWAYS SOLUBLE, WILL NOT BE CONSIDERED A REACTANT IN THE MIXTURES**